

# Download File PDF Chemistry Workbook Chapter 15 Water And Aqueous Systems Answers

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Section 15.1  
Water and its Properties

**OBJECTIVES:**

- Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding.
- Describe the structure of ice.

**The Water Molecule: a Review**

- Water is a simple tri-atomic molecule,  $H_2O$ .
- Each O-H bond is \_\_\_\_\_, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values).
- bond angle of water =  $105^\circ$
- due to the \_\_\_\_\_, the O-H bond polarities do not cancel. This means \_\_\_\_\_.
- Water's **heat capacity** and **ability to hydrogen bond** gives it many special properties!
- Water molecules are attracted to one another by dipole interactions.
- This hydrogen bonding gives water:  
a) its \_\_\_\_\_ and  
b) its \_\_\_\_\_.

**a) High Surface Tension?**

- liquid water acts like it has a "skin"  
— glass of water bulges over the top
- Water forms round drops  
— spray water on grassy surface
- All because water hydrogen bonds.

**Surface Tension**

- One water molecule can hydrogen bond to another because of this electrostatic attraction.
- Also, hydrogen bonding occurs with many other molecules surrounding them on all sides.

A water molecule in the middle of a solution is pulled in all directions.

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